AQA

Please write clearly in block capitals.

Centre number | | | | | | Candidate number | | | | |

Surname _____

Forename(s) _____

Candidate signature _____

I declare this is my own work.

A-level CHEMISTRY

Paper 3

Friday 21 June 2024 Morning Time allowed: 2 hours

Materials

For this paper you must have:
- the Periodic Table/Data Booklet, provided as an insert (enclosed)
- a ruler with millimetre measurements
- a scientific calculator, which you are expected to use where appropriate.

Instructions
- Use black ink or black ball-point pen.
- Fill in the boxes at the top of this page.
- Answer **all** questions.
- You must answer the questions in the spaces provided. Do **not** write outside the box around each page or on blank pages.
- If you need extra space for your answer(s), use the lined pages at the end of this book. Write the question number against your answer(s).
- All working must be shown.
- Do all rough work in this book. Cross through any work you do not want to be marked.

For Examiner's Use	
Question	Mark
1	
2	
3	
4	
5	
Section B	
TOTAL	

Information
- The marks for questions are shown in brackets.
- The maximum mark for this paper is 90.

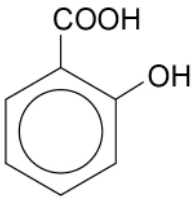
Advice
- You are advised to spend 70 minutes on **Section A** and 50 minutes on **Section B**.

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Section A

Answer **all** questions in this section.

0 1 The structure of 2-hydroxybenzenecarboxylic acid is shown.



0 1 . 1 Give the equation for the reaction of 2-hydroxybenzenecarboxylic acid with methanol.

In your equation, include the **skeletal** formula of the organic product.

[2 marks]

Aspirin is produced from 2-hydroxybenzenecarboxylic acid by reaction with ethanoic anhydride in the presence of concentrated phosphoric acid.

Method

1. Add 2-hydroxybenzenecarboxylic acid to a conical flask.
2. Add excess ethanoic anhydride.
3. Add a few drops of concentrated phosphoric acid.
4. Heat the flask to 85 °C for 10 minutes.
5. Cool the flask and pour the contents into 150 cm^3 of cold water.
6. Filter and wash the impure solid aspirin.
7. Recrystallise the aspirin using a 50:50 mixture of water and ethanol.
8. Check the purity of the aspirin.

0 1 . 2 Aspirin can also be produced by reacting 2-hydroxybenzenecarboxylic acid with ethanoyl chloride.

State why ethanoic anhydride is preferred to ethanoyl chloride for this preparation.

[1 mark]

Do not write outside the box

0 1 . 3 Give the name of the mechanism for the reaction of 2-hydroxybenzenecarboxylic acid with ethanoic anhydride.

[1 mark]

0 1 . 4 Suggest the role of the concentrated phosphoric acid.

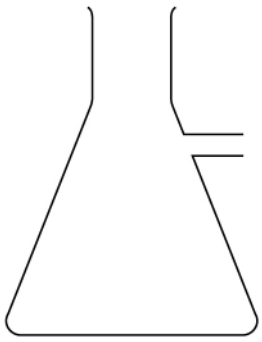
[1 mark]

0 1 . 5 Suggest why reflux is **not** essential when the flask is heated to 85 °C for 10 minutes.

[1 mark]

0 1 . 6 Complete and label the diagram to show how the impure solid is filtered.

[2 marks]



Question 1 continues on the next page

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0 1 . 7 Suggest the identity of **two** impurities present in the filtered solid aspirin before it is washed in Step 6 of the method.

[2 marks]

Impurity 1 _____

Impurity 2 _____

0 1 . 8 Describe the recrystallisation process used to purify the aspirin.

Include the method and apparatus used.

[6 marks]

Do not write outside the box

0 1 . 9 State the physical property that is measured to check the purity of the aspirin.

Describe **two** ways the result would show that the product is impure.

[3 marks]

Physical property _____

1 _____

2 _____

19

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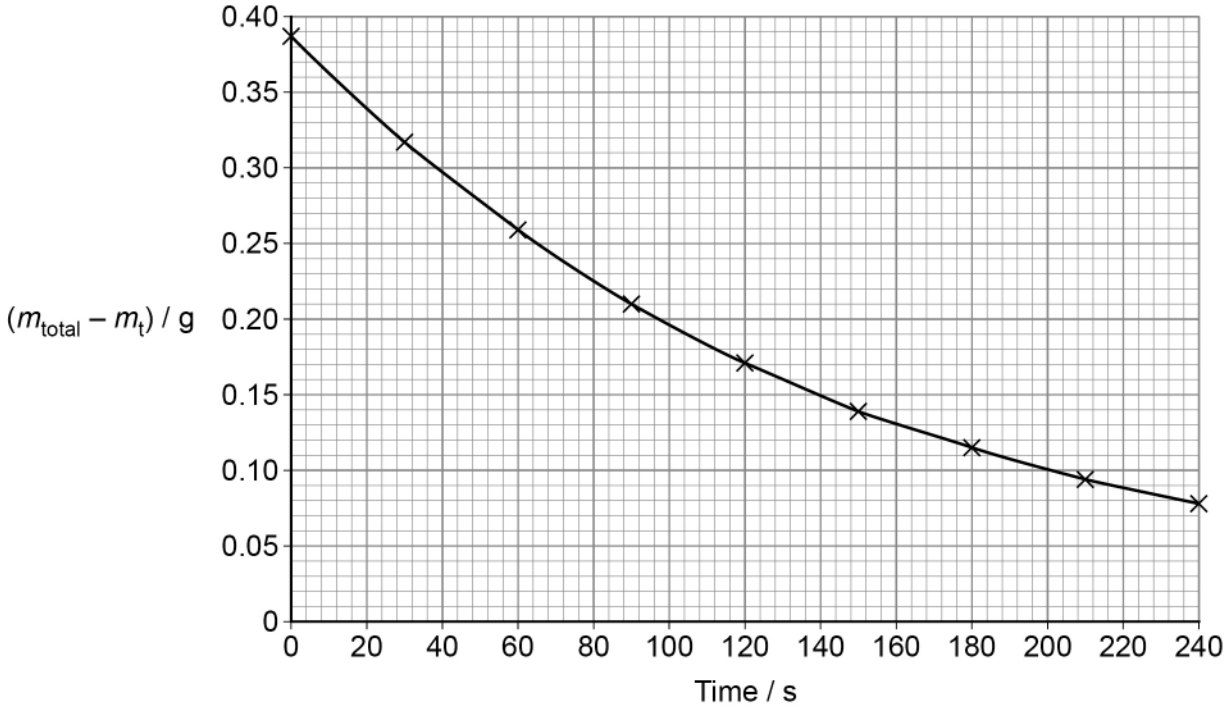
0 2 The rate of reaction between calcium carbonate and hydrochloric acid is investigated using a continuous monitoring method.

Method

- Place a conical flask on a balance and add approximately 20 g of large marble chips.
- Add 50 cm^3 of 0.4 mol dm^{-3} hydrochloric acid.
- Place a loose cotton wool plug in the neck of the flask.
- Zero the mass reading on the balance.
- Start a timer.
- Record the loss in mass (m_t) every 30 seconds for 4 minutes.
- Wait for the reaction to finish and record the total mass loss (m_{total}).
- Plot a graph of ($m_{total} - m_t$) against time.

Figure 1 shows a graph of the results obtained during the first 240 s

Figure 1



0 2 . 1 Suggest why a loose cotton wool plug is placed in the neck of the flask, instead of leaving the flask open or inserting a bung.

[2 marks]

Instead of leaving the flask open _____

_____

Instead of inserting a bung _____

_____

Do not write outside the box

0 2 . 2 20 g of large marble chips is a large excess of calcium carbonate.

Suggest why using a large excess of calcium carbonate means that the rate is only affected by the changing concentration of the hydrochloric acid.

[1 mark]

0 2 . 3 The mass of carbon dioxide produced in time t is equal to m_t.

The total mass of CO_2 produced during the reaction is equal to m_{total}.

Explain why ($m_{total} - m_t$) is proportional to the concentration of hydrochloric acid remaining in the flask at time t.

[2 marks]

Question 2 continues on the next page

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0 2 . 4 **Table 1** shows the rate of reaction, calculated from the gradient of the curve, at five different times.

($m_{total} - m_t$) is proportional to the concentration of unreacted HCl at time *t*.

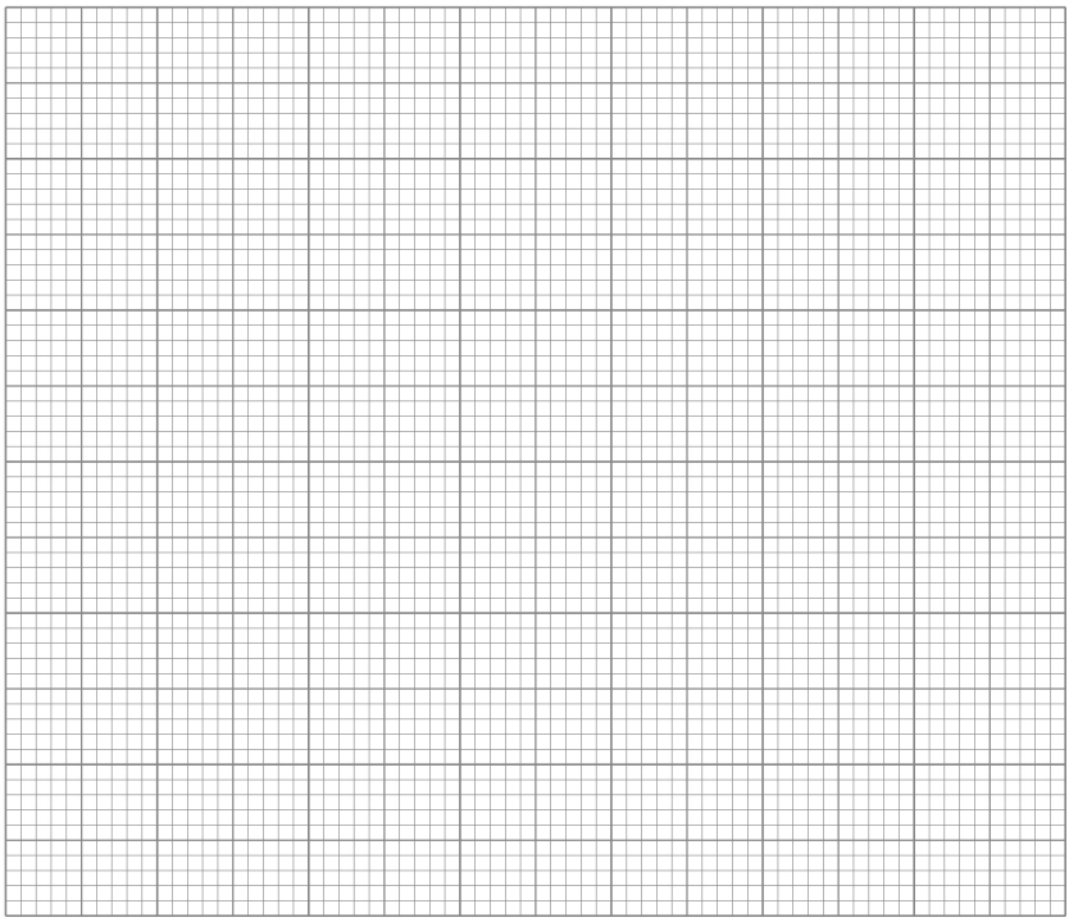
Table 1

Rate of reaction / g s^{-1}	23.0×10^{-4}	19.0×10^{-4}	15.7×10^{-4}	11.5×10^{-4}	6.67×10^{-4}
($m_{total} - m_t$) / g	0.340	0.280	0.225	0.170	0.100

On the grid in **Figure 2** plot the rate of reaction (*y*-axis) against ($m_{total} - m_t$) (*x*-axis).

[3 marks]

Figure 2



Do not write outside the box

0 2 . 5 State how the graph in **Figure 2** confirms that the rate equation for this reaction is

$$Rate = k[HCl]$$

[1 mark]

0 2 . 6 In this experiment the variable measured is mass loss.

The rate of this reaction at a constant temperature can be investigated in other ways.

Suggest **two** other variables that can be measured instead of mass loss.

[2 marks]

1 _____

2 _____

11

Turn over for the next question

Do not write outside the box

0 3 The thermal decomposition of but-3-en-1-ol is investigated at different temperatures (T).

$$CH_2{=}CHCH_2CH_2OH \rightarrow CH_2{=}CHCH_3 + HCHO$$

The results from the investigation are used to calculate the rate constant, k, at each temperature.

Table 2 shows some of the results.

Table 2

T / K	$\frac{1}{T}$ / K^{-1}	k / s^{-1}	ln k
553	1.81×10^{-3}	4.6×10^{-4}	–7.68
563	1.78×10^{-3}	8.4×10^{-4}	–7.08
573		15.6×10^{-4}	
583	1.72×10^{-3}	28.0×10^{-4}	–5.88
593	1.69×10^{-3}	49.9×10^{-4}	–5.30

0 3 . 1 Complete **Table 2** with the missing values at 573 K

[1 mark]

0 3 . 2 The overall order of the reaction can be deduced from a piece of information in one of the column headings in **Table 2**.

Identify this piece of information and deduce the overall order.

[2 marks]

Piece of information _____

Overall order _____

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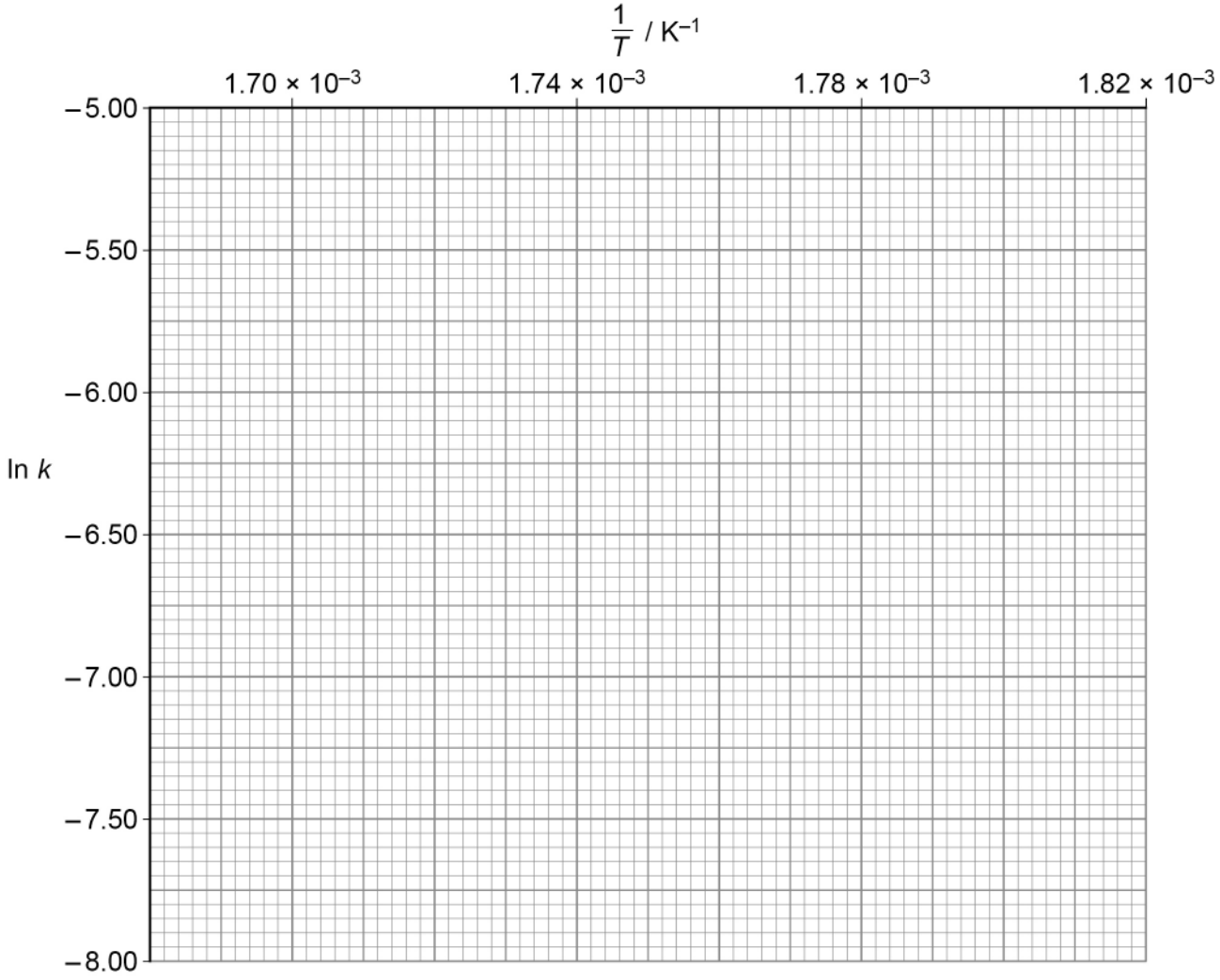
0 3 . 3 The Arrhenius equation can be written in the form shown.

$$\ln k = \ln A - \frac{E_a}{RT}$$

On the grid in **Figure 3** plot a graph of ln k against $\frac{1}{T}$

[2 marks]

Figure 3



Question 3 continues on the next page

Do not write outside the box

0 3 . 4 Use your graph from Question **03.3** to calculate a value for E_a, in kJ mol^{-1}, for the thermal decomposition of but-3-en-1-ol.

The gas constant, R = 8.31 J K^{-1} mol^{-1}

[3 marks]

E_a _____ kJ mol^{-1}

0 3 . 5 2-Methylpent-4-en-2-ol decomposes in a similar way to but-3-en-1-ol, to produce an alkene and a carbonyl compound.

Deduce the structures of the alkene and the carbonyl compound.

[2 marks]

$CH_2{=}CHCH_2C(CH_3)_2OH \rightarrow$ _____ + _____

alkene carbonyl compound

10

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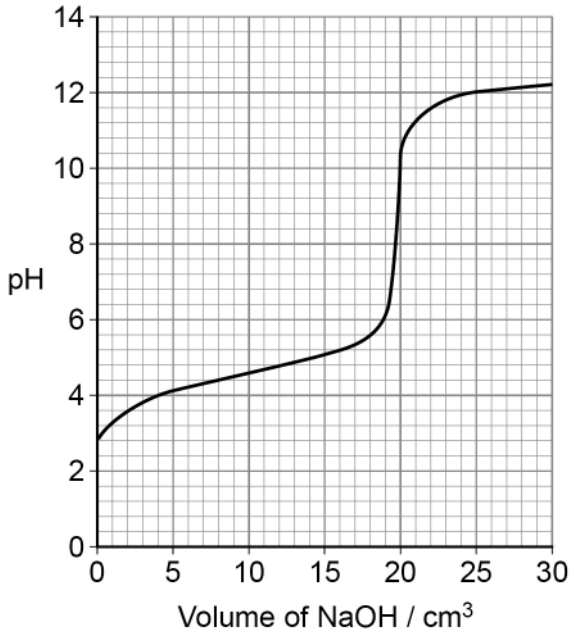
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ANSWER IN THE SPACES PROVIDED

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0 4 **Figure 4** shows how the pH changes as 0.100 mol dm^{-3} sodium hydroxide solution is added to 25.0 cm^3 of 0.0800 mol dm^{-3} aqueous propanoic acid at 298 K

Figure 4



0 4 . 1 Propanoic acid is a weak acid.

State the meaning of weak in this context.

[1 mark]

0 4 . 2 Suggest why a student doing an experiment to produce the curve in **Figure 4** would add the sodium hydroxide solution dropwise around the equivalence point.

[1 mark]

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0 4 . 3 Give an expression for K_a for propanoic acid (CH_3CH_2COOH).

Use this expression to show that pH = pK_a when half of the propanoic acid has reacted with sodium hydroxide.

[3 marks]

K_a

0 4 . 4 Use the pH from **Figure 4**, when half of the propanoic acid has reacted, to calculate K_a at 298 K

[2 marks]

K_a _____ mol dm^{-3}

Question 4 continues on the next page

Turn over ►

Do not write outside the box

0 4 . 5 When sodium hydroxide solution is added to aqueous propanoic acid, the solution formed acts as a buffer when between 5 cm^3 and 15 cm^3 have been added.

Explain why the pH stays approximately constant during this part of the experiment.

[2 marks]

0 4 . 6 Methyl orange and universal indicator are **not** suitable indicators for the titration of solutions of propanoic acid with sodium hydroxide.

State the reason why each indicator is **not** suitable.

[2 marks]

Methyl orange

Universal indicator

11

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0 5 . 1 Some complexes containing transition metal ions are coloured.

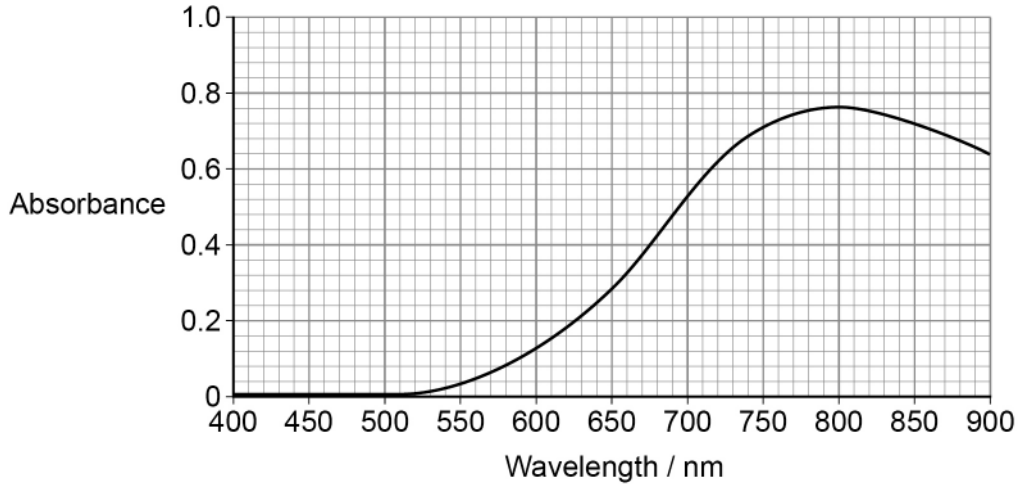
- Explain why some complexes containing transition metal ions are coloured.
- List the factors that affect the colour.
- Describe how colorimetry can be used to determine the concentration of a coloured complex.

[6 marks]

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0 5 . 2 **Figure 5** shows the visible spectrum of $[Cu(H_2O)_6]^{2+}$

Figure 5



Use the wavelength at the peak of the curve in **Figure 5** to calculate the change in energy, in J, of an electron when it absorbs radiation with this wavelength.

the Planck constant, $h = 6.63 \times 10^{-34}$ J s

speed of light, $c = 3.00 \times 10^{8}$ m s^{-1}

[3 marks]

Change in energy _____ J

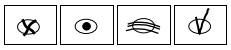
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Section B

Answer **all** questions in this section.

Only **one** answer per question is allowed.
For each question completely fill in the circle alongside the appropriate answer.

CORRECT METHOD  WRONG METHODS 

If you want to change your answer you must cross out your original answer as shown. 

If you wish to return to an answer previously crossed out, ring the answer you now wish to select as shown. 

You may do your working in the blank space around each question but this will not be marked. Do **not** use additional sheets for this working.

0 6 Which row contains two species with different numbers of electrons?

[1 mark]

A NH_3 and HF

B CO_3^{2-} and NO_3^-

C H_3O^+ and HF_2^+

D CH_4 and NH_2^-

0 7 Which element has the highest third ionisation energy?

[1 mark]

A Li

B Be

C K

D Ca

Do not write outside the box

0 8 Compound **P** is converted into compound **R** by a two-stage synthesis via compound **Q**.

The yields for the individual steps are:

P → **Q** 50%
Q → **R** 30%

What is the overall yield of **R** in this synthesis?

[1 mark]

A 15% ⬭

B 30% ⬭

C 40% ⬭

D 80% ⬭

0 9 What is the formula of ammonium ethanedioate?

[1 mark]

A $NH_4(C_2O_4)_2$ ⬭

B $(NH_4)_2C_2O_4$ ⬭

C $NH_4C_2O_4$ ⬭

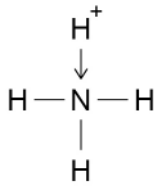
D $(NH_4)_2(C_2O_4)_3$ ⬭

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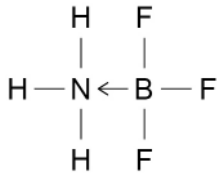
1 0 Which diagram shows the formation of a dative covalent bond?

[1 mark]

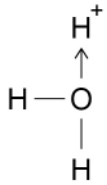
A



B



C



D



1 1 Which substance does **not** have any bond angles of 120°?

[1 mark]

A benzene

B boron trifluoride

C cyclohexane

D graphite

Do not write outside the box

1 2 Which molecule does **not** have a permanent dipole?

[1 mark]

A NH_3 ⬭

B PCl_3 ⬭

C SCl_2 ⬭

D $SiCl_4$ ⬭

1 3 Which compound forms the greatest number of hydrogen bonds per molecule in the liquid state?

[1 mark]

A CH_3CH_2COOH ⬭

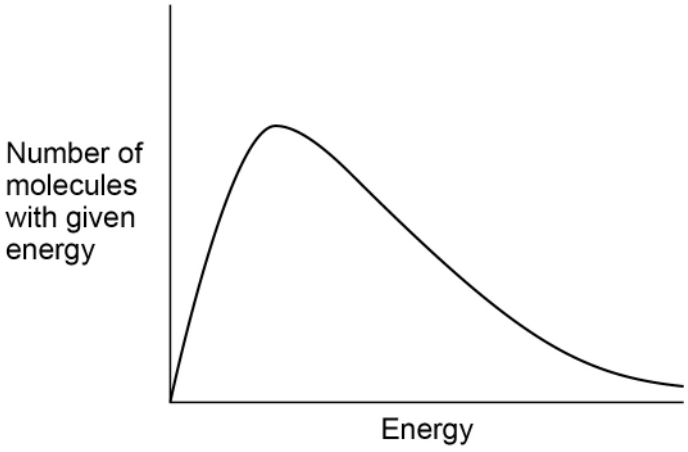
B $CH_3CH_2OCH_3$ ⬭

C CH_3CH_2CHO ⬭

D $CH_3CH_2CH_2OH$ ⬭

Do not write outside the box

1 4 The Maxwell–Boltzmann distribution curve is shown for a gaseous reaction mixture.



What is represented by the total area under the curve?

[1 mark]

A The total energy of the molecules in the reaction mixture ⬭

B The total energy of reacting molecules in the reaction mixture ⬭

C The total number of molecules in the reaction mixture ⬭

D The total number of reacting molecules in the reaction mixture ⬭

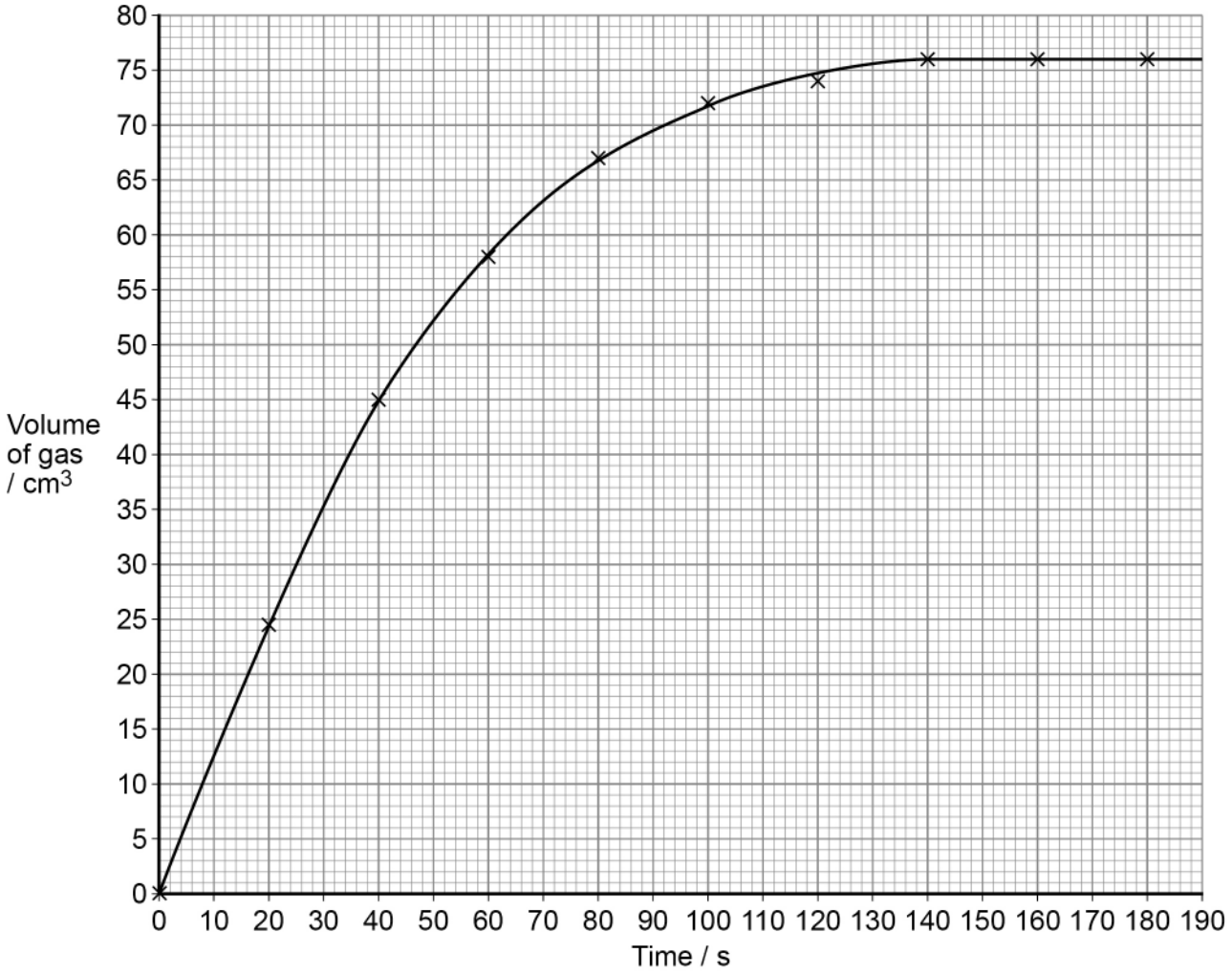
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1 5 A student investigated the reaction between magnesium and hydrochloric acid.

$$Mg(s) + 2HCl(aq) \longrightarrow MgCl_2(aq) + H_2(g)$$

The results are plotted on this graph.



Which value is closest to the rate of reaction, in $cm^3\ s^{-1}$, at 70 s?

[1 mark]

A 0.4 ⬭

B 0.9 ⬭

C 1.1 ⬭

D 2.5 ⬭

Do not write outside the box

1 6 Which statement explains why the rate of a reaction increases when a catalyst is added at a constant temperature?

[1 mark]

A The collision frequency increases because there is a decrease in activation energy. ⬭

B The collision frequency increases because there is an increase in the average energy of the particles. ⬭

C The proportion of successful collisions increases because there is a decrease in activation energy. ⬭

D The proportion of successful collisions increases because there is an increase in the average energy of the particles. ⬭

1 7 A cell with EMF = +2.15 V is made from two electrodes.

The half-equations for the two electrodes are shown.

positive electrode:

$$PbO_2(s) + 3H^+(aq) + HSO_4^-(aq) + 2e^- \longrightarrow PbSO_4(s) + 2H_2O(l)$$

negative electrode:

$$PbSO_4(s) + H^+(aq) + 2e^- \longrightarrow Pb(s) + HSO_4^-(aq) \quad E^\ominus = -0.46\ V$$

What is the standard electrode potential of the PbO_2 / $PbSO_4$ electrode?

[1 mark]

A –2.61 V ⬭

B –1.69 V ⬭

C +1.69 V ⬭

D +2.61 V ⬭

Turn over ►

Do not write outside the box

1 8 Values of the ionic product of water (K_w) at different temperatures are given.

$K_w = 6.40 \times 10^{-15}$ mol^2 dm^{-6} at 18 °C
$K_w = 1.00 \times 10^{-14}$ mol^2 dm^{-6} at 25 °C

Which statement is correct?

[1 mark]

A The concentration of hydroxide ions in water at 18 °C is 8.00×10^{-8} mol dm^{-3} ⬭

B The dissociation of water into ions is an exothermic process. ⬭

C The pH of water is the same at 25 °C and at 18 °C ⬭

D Water becomes less acidic as the temperature is raised. ⬭

1 9 Consider the Period 3 elements from sodium to chlorine.

Which statement is correct?

[1 mark]

A Sodium has the smallest atomic radius. ⬭

B Aluminium has the highest melting point. ⬭

C Sulfur is the most electronegative. ⬭

D Chlorine has the highest first ionisation energy. ⬭

2 0 Which statement correctly describes a trend down Group 7 from Cl to I?

X represents Cl, Br or I

[1 mark]

A The boiling point of HX increases. ⬭

B The bond dissociation energy of H–X increases. ⬭

C The standard electrode potential value for $X_2(aq) + 2e^- \rightarrow 2X^-(aq)$ becomes more positive. ⬭

D The solubility of AgX in ammonia increases. ⬭

Do not write outside the box

2 1 Which statement about chloride ions is correct?

[1 mark]

A They are oxidised by concentrated sulfuric acid. ⬭

B They form a cream precipitate with silver nitrate solution. ⬭

C They form a cobalt(II) complex with a tetrahedral shape. ⬭

D They have the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^4$ ⬭

2 2 Aqueous aluminium sulfate is added to aqueous sodium carbonate.

What are the formulas of the precipitate and the gas formed?

[1 mark]

A $Al_2(CO_3)_3$ and SO_2 ⬭

B $Al_2(CO_3)_3$ and CO_2 ⬭

C $Al(H_2O)_3(OH)_3$ and SO_2 ⬭

D $Al(H_2O)_3(OH)_3$ and CO_2 ⬭

2 3 What is the correct structural formula for 4-chloro-2-methylpent-2-enoic acid?

[1 mark]

A $CH_3CCl{=}CHCH(CH_3)COOH$ ⬭

B $(CH_3)_2C{=}CHCHClCOOH$ ⬭

C $CH_3CHClCH{=}C(CH_3)COOH$ ⬭

D $(CH_3)_2CHCH{=}CClCOOH$ ⬭

Do not write outside the box

2 4 In which conversion does a nucleophile attack the organic reactant?

[1 mark]

A $CH_3CH_2CH_3 \longrightarrow CH_3CHClCH_3$ ⬭

B $CH_3CH{=}CH_2 \longrightarrow CH_3CHBrCH_3$ ⬭

C $CH_3CH_2CH_2OH \longrightarrow CH_3CH{=}CH_2$ ⬭

D $CH_3CH_2CHO \longrightarrow CH_3CH_2CH(OH)CN$ ⬭

2 5 Compound **Y** has the structural formula $CH_3COOCH_2CH(CH_3)_2$

Which compound is a position isomer of **Y**?

[1 mark]

A 5-hydroxyhexan-3-one ⬭

B butyl ethanoate ⬭

C hexanoic acid ⬭

D propyl propanoate ⬭

2 6 Which compound shows *E*–*Z* isomerism?

[1 mark]

A 2,3-dimethylbut-1-ene ⬭

B 2,3-dimethylbut-2-ene ⬭

C 2-methylpent-2-ene ⬭

D 3-methylpent-2-ene ⬭

Do not write outside the box

2 7 What is the product when 3-methylbutan-2-one reacts with acidified KCN?

[1 mark]

A 2-hydroxy-2,3-dimethylbutanenitrile ⬭

B 3-hydroxy-2,3-dimethylbutanenitrile ⬭

C 2-hydroxy-3-methylpentanenitrile ⬭

D 3-hydroxy-2-methylpentanenitrile ⬭

2 8 Which statement concerning nylon-6,6 is correct?

[1 mark]

A Butanedioic acid is one of the reactants used to make nylon-6,6 ⬭

B Nylon-6,6 is an addition polymer. ⬭

C Nylon-6,6 can be hydrolysed by aqueous sodium hydroxide. ⬭

D All molecules of nylon-6,6 have the same relative molecular mass. ⬭

2 9 Which statement about the industrial production of ethanol from ethene at 300 °C is correct?

$$C_2H_4(g) + H_2O(g) \rightleftharpoons C_2H_5OH(g) \qquad \Delta H = -46 \text{ kJ mol}^{-1}$$

[1 mark]

A The use of an acid catalyst increases the yield of ethanol. ⬭

B The reaction is slower than fermentation. ⬭

C An increase in temperature, at constant pressure, increases the value of K_p. ⬭

D An increase in pressure, at constant temperature, increases the equilibrium yield of ethanol. ⬭

Turn over ►

Do not write outside the box

3 0 Which compound is formed by the acid hydrolysis of phenyl benzenecarboxylate?

[1 mark]

A $C_6H_5CH_2OH$ ⬭

B C_6H_5CHO ⬭

C $C_6H_5COCH_3$ ⬭

D C_6H_5COOH ⬭

3 1 Which type of polymer is most difficult to hydrolyse?

[1 mark]

A polyalkene ⬭

B polyamide ⬭

C polyester ⬭

D protein ⬭

3 2 In which polymer does hydrogen bonding occur between the polymer chains?

[1 mark]

A a polyalkene ⬭

B a polyamide ⬭

C a polychloroalkene ⬭

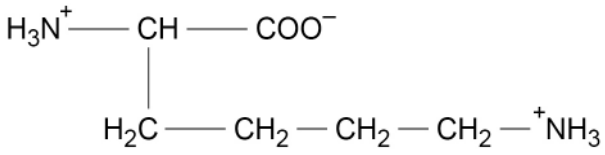
D a polyester ⬭

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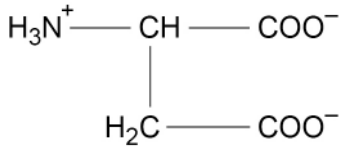
3 3 Which structure shows a zwitterion of an amino acid?

[1 mark]

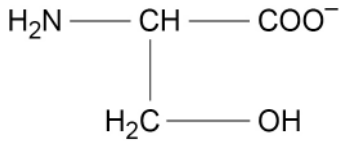
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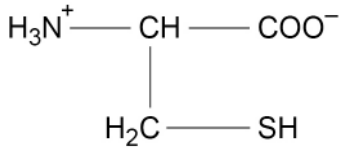
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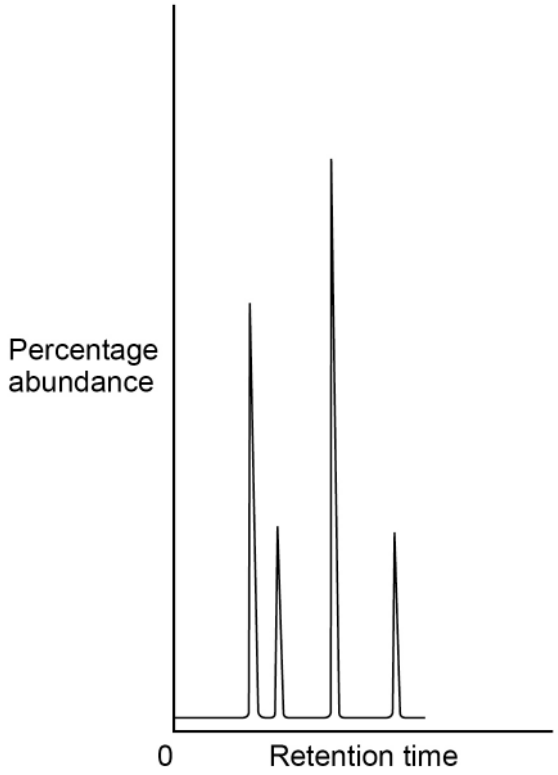


Turn over for the next question

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3 4 The diagram shows a gas chromatogram for a sample containing four isomers with the molecular formula $C_6H_{12}O_2$

The carrier gas is nitrogen and the stationary phase is polar.

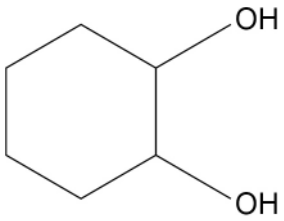


Which of the four isomers in this sample is the most abundant?

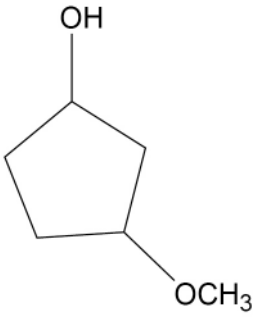
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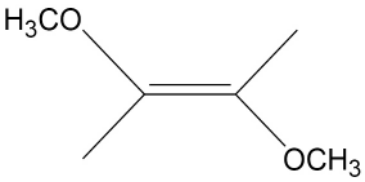
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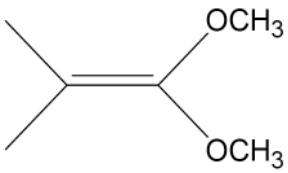
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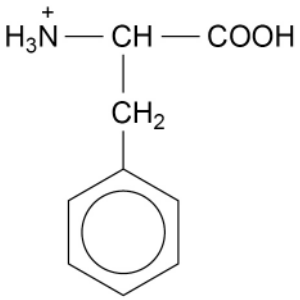
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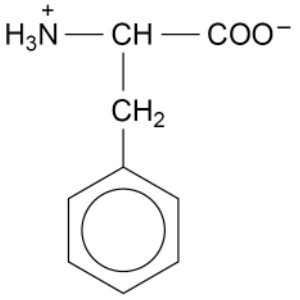
3 5 Which structure is formed by phenylalanine in solution at pH = 3?

[1 mark]

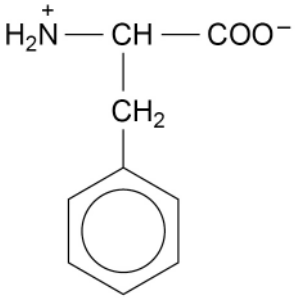
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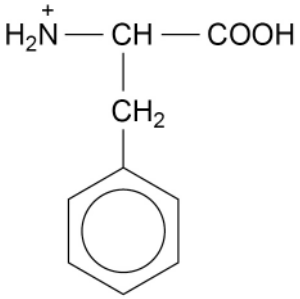
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END OF QUESTIONS

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